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An acid is a substance that  
contains hydrogen and ionizes to  
produce hydrogen ions in

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aqueous solution; A base is a substance that contains a hydroxide group and dissociates to produce a hydroxide ion in aqueous solution.

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## Section 19.1 – Acid-Base Theories

Acids have a sour taste, change the color of an indicator, can be strong or weak electrolytes in aqueous solution, and react with metals.

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states that an acid is a hydrogen ion donor, and a base is a hydrogen ion acceptor. It is a more inclusive model of acids and bases. conjugate acid. is the species produced when a base accepts a hydrogen ion.



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Conjugate base. is the species produces when a acid donates a hydrogen ion. conjugate acid-base pair.

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according to Arrhenius, acids are

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hydrogen-containing compounds that ionise to yield hydrogen ions ( $H^+$ ) in aqueous solution. bases are compounds that ionise to yield hydroxide ions ( $OH^-$ ) in aqueous solution.

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Description. Acids/Bases

descriptions & formulas with pH.

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Chapter 19 Acids and Bases Now, consider the equation for the

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ionization of hydrogen fluoride in water. According to the Brønsted-Lowry definition, the equation is written this way. Agricultural Technician

$$\text{HF} + \text{H}_2\text{O} \rightleftharpoons \text{H}_3\text{O}^+ + \text{F}^-$$

What are the conjugate acid-base pairs?

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Salts. STUDY. PLAY. Tastes sour.  
Acid. Changes the color of an acid-  
base indicator (acid, base, or  
both) Both. Can be strong or

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weak electrolytes in aqueous solution. Acid. In drinks, citrus, pop, etc. Acid. This acid is associated with protein. amino acid. Used in fragrances and flavors.

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Arrhenius Acids. Acids are hydrogen containing compounds, ionize to yield hydrogen ions in aqueous solutions. Bases are compounds that ionize to yield hydroxide ions in aqueous

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**Solutions: Monoprotic acids.** acids that contain 1 ionizable hydrogen such as nitric acid ( $\text{HNO}_3$ ) diprotic acids.

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Brønsted-Lowry acids and bases  
come in pairs. A conjugate base is  
the remainder of the B-L acid,

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after it donates its  $H^+$ . A conjugate acid is the substance formed when the B-L base gains a  $H^+$ . Thus, a conjugate acid-base pair is related by the . loss or gain. of a single hydrogen ion.

*Chapter 19 Acids, Bases, and*

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If the salt is from a strong acid and base then the pH is 7, for example  $\text{KNO}_3$ . 3. A salt formed between a strong acid and a conjugate of a weak base is an acid salt, example  $\text{NH}_4\text{Cl}$ , pH = acidic. ... Chapter 19 Acids,

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Bases, and Salts Last modified by:  
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Chapter 19 Acids, Bases, and  
Salts. Acids and Bases: Basics.  
These are really just a specific



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type of chemical compound. No mysteries here. 1) They MUST be ionic compounds, and. MUST dissolve in water, that's the only way to be chemically active. We call this. disassociation.

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Salts". Use these activities to study the vocabulary and major concepts presented in this chapter. compounds derived from the reaction of a strong base with a weak acid or a strong acid with a weak base or a solution of a weak base and one of its salts.

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#46 - equilibrium constants

Chapter 19, Slides 1-4: eReview  
acid naming Chapter 19, Slides

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19, Slides 8-9: Properties of Bases  
Chapter 19, Slides 10-12:  
Definitions of Acids Chapter 19,  
Slides 13-15: pH, calculations,  
examples

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acids and bases? acids taste sour,  
will change the color of an acid-  
base indicator, and can be strong  
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